Name: _____ Date: _____

Neutrons

1. Neutrons are neutral.
   a. True
   b. False

2. Neutrons are found in the rings around the Bohr-Rutherford diagram.
   a. True
   b. False

3. Where is the neutron located?
   a. outside the nucleus
   b. in the nucleus

4. Uranium-238 contains 146 neutrons.
   a. True
   b. False

5. To find the number of neutrons in an atom, you subtract the number of protons from the _____.
   a. atomic number
   b. electrons
   c. mass number
   d. atomic symbol

6. The mass of a neutron is
   a. about equal to the mass of a proton.
   b. much less than the mass of a proton.
   c. about equal to the mass of an electron.
   d. much less than the mass of an electron.

7. An atom of fluorine has an atomic number of 9 and a mass number of 19. How many neutrons does fluorine have?
   a. 9
   b. 10
   c. 19
   d. 29

8. In general, as the atomic numbers increase on the Periodic Table, the number of neutrons in the atoms _____.
   a. decrease
   b. increase
   c. fluctuate
   d. remain the same

9. Which element has 4 neutrons?
   a. an element with an atomic number of 2 and an atomic mass of 4
   b. an element with an atomic number of 3 and an atomic mass of 7
   c. an element with an atomic number of 4 and an atomic mass of 9
   d. an element with an atomic number of 7 and an atomic mass of 14

10. What is unique about the number of neutrons in a normal hydrogen atom?
    a. It does not have any neutrons.
    b. It has more neutrons than protons.
    c. It has an equal number of neutrons and protons.
    d. It has an equal number of neutrons and electrons.